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Chemical bond. a mutual electrical attraction between the nuclei and valance electrons of different atoms that binds the atoms together.

Chemical formula. a formula that indicates the relative numbers of

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atoms of each kind in a chemical compound by using atomic symbols and numerical subscripts.

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Bond energy is therefore the energy required to break a chemical bond. State the octet rule (6.2) Chemical compounds tend to form so that each atom, by gaining,

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losing, or sharing electrons, has an octet of electrons in its highest (occupied) energy level.

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chemical bond. a mutual electrical attraction between the nuclei and valence electrons of different atoms that binds the atoms together. independent particles. as \_\_\_\_\_, atoms have a tremendous amount of potential energy and

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nature favors

arrangements which

reduce potential

energy.

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Bond energy is the energy required to break a chemical bond and form neutral isolated atoms. The bond energy for a hydrogen-hydrogen bond is 436 kJ/mol. In other words, it takes 436 kJ of energy to break the bonds in a mole of hydrogen molecules. **READING CHECK 3.** Define the

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following terms in your own words.

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Section 6.2 – Covalent Bonding. A covalent bond is a chemical bond in which two atoms share a pair of valence electrons.

When two atoms share one pair of electrons, the bond is called a single bond.

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## **Chapter 6 - Chemical**

## **Bonds**

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1.

a A chemical bond between atoms results from the attraction between the valence electrons and of

different atoms. (a) nuclei (c) isotopes (b) inner electrons (d)

Lewis structures 2. b A

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covalent bond consists of (a) a shared electron.

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Know the following definitions: molecule, molecular compound, chemical formula, molecular formula, crystal= oppositely charged ions form a crystal like shape where electrons of adjacent ions repel

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each other, resonance, octet, bond length, covalent bonds, ionic bonds, lattice energy=  
The energy released when 1 mol of an ionic crystalline compound is formed from gaseous ions.

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## **Chemistry Chapter No # 6 Chemical Bonding**

Chapter 6 - Chemical Bonding In previous chapters, students have studied subatomic particles and the properties of individual atoms. In Chapter 6, we will begin studying how

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atoms interact with each other to form chemical bonds.

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notes: 1. In \_\_\_\_\_ the  
electron pairs in the



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central atom are directed at  $109^\circ$  angle. (a)  $\text{H}_2\text{O}$  (b)  $\text{CO}_2$  (c)  $\text{CCl}_4$  (d)  $\text{SO}_2$ . 2. According to electron pair repulsion theory shape of the molecule depends on \_\_\_\_.

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SECTION 6-1 SHORT

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ANSWER Answer the following questions in the space provided. 1. A chemical bond between atoms results from the ... [Filename: HC2SR061.PDF] - Read File Online - Report Abuse

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force that holds different atoms in a molecule is called chemical bond. • Octet Rule. Atoms of different elements take part in chemical combination in order to complete their octet or to attain the noble gas configuration.

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and Molecular  
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**XI**

Chemical bonds are the glue that hold molecules together. We will learn about the different kinds of bonds, ways chemists draw bonds and molecules, and how the type of chemical bonding affects the bulk properties of a material. We will cover

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electronegativity,  
Lewis dot structures,  
VSEPR, bond  
hybridization, and  
ionic, covalent, and  
metallic bonds.

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